Due Date:

Preview of Section 6.1 “Compounds and Molecules” Name:_____

Define the following vocabulary:

- Chemical Bond
- Chemical Structure
- Bond Length
- Bond Angle

Write down 2 things from the following sections:

- Chemical Bonds

- Chemical Structure
    - Some models represent bond lengths and angles.
    - Space-filling models show the space occupied by atoms.
    - Bonds can bend, stretch, and rotate without breaking.

- How Does Structure Affect Properties
    - Compounds with network structures are strong solids.

Due Date:

- Some networks are made of bonded ions.

- Some compounds are made of molecules.

- The strength of attractions between molecules varies.

Due Date:

Preview of Section 6.2 “Ionic and Covalent Bonding” Name:_____

Define the following vocabulary:

- Ionic Bond
- Covalent Bond
- Metallic Bond
- Polyatomic Ion

Write down 2 things from the following sections:

- Why Do Chemical Bonds Form?

- Ionic Bonds
    - Ionic bonds are formed by the transfer of electrons.
    - Ionic compounds are in the form of networks, not molecules.
    - When melted or dissolved in water, ionic compounds conduct electricity.

- Covalent Bonds
    - Atoms may share more than one pair of electrons.
    - Atoms do not always share electrons equally.

Due Date:

- Metallic Bonds
    - Electrons move freely between metal atoms.
- Polyatomic Ions
    - There are many common polyatomic ions.
    - Parentheses group the atoms of a polyatomic ion.
    - Some names of polyatomic anions relate to the oxygen content of the anion.

Due Date:

Preview of Section 6.3 “Compound Names and Formula” Name:_____

Define the following vocabulary:

- Alkali Metal
- Alkaline–Earth Metal
- Transition Metal
- Noble Gas
- Halogen

Write down 2 things from the following sections:

- Classifying Elements Further

- Metals

    - The alkali metals are very reactive.

    - Alkaline–earth metals form compounds that are formed in limestone and in the human body.

    - Transition metals are in the middle of the periodic table.

    - Some elements are synthetic.

- Nonmetals

Due Date:

- The noble gases are relatively inert.
- The halogens combine easily with metals to form salts.
- Nonmetals and their compounds are plentiful on Earth.
- Carbon can form many compounds.

- Semiconductors
  - Hydrogen is in a class by itself.